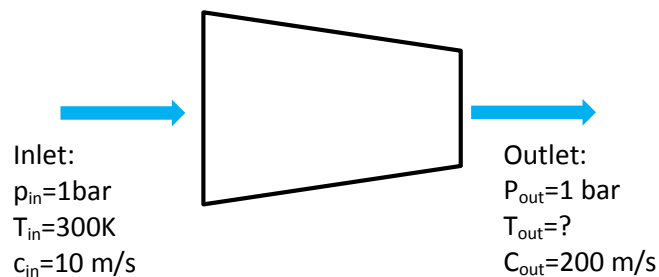


Air flows through an adiabatic and horizontal placed pipe with changeable cross section. At the inlet side air flows with velocity $c_{in}=10$ m/s under the pressure $p_{in}=1$ bar and the temperature $\vartheta_{in}=25^{\circ}\text{C}$. At the outlet side air flows out of the pipe with velocity $c_{out}=150$ m/s under the pressure $p_{in}=0.8$ bar.

- 1) Determine whether this pipe is nozzle or diffuser without calculating the ratio of cross section of outlet and inlet A_{out}/A_{in}
- 2) Determine now the ratio of cross section of outlet and inlet A_{out}/A_{in}

Air can be considered as perfect gas with $c_p=1$ J/K·g and $R=0.29$ J/K·g

Solution:



1)

The 1.Law of thermodynamics:

$$\sum_i \dot{Q}_i + \sum_j \dot{W}_j + \sum_k \dot{m}_k \cdot \left(h + \frac{c^2}{2} + g \cdot z \right)_k = \frac{d}{d\tau} \left(\sum_l m_l \cdot (\bar{u} + \bar{e}_{kin} + \bar{e}_{pot})_l \right)$$

Stationary $\rightarrow d/d\tau=0$

Adiabatic $\rightarrow \sum Q=0$

Without input and output (work) $\rightarrow \sum W=0$

Horizontal placed \rightarrow no change in potential energy $\rightarrow \Delta(g \cdot z)=0$

Then we can obtain the simplified 1.Law:

$$h_{in} + \frac{c_{in}^2}{2} = h_{out} + \frac{c_{out}^2}{2}$$

wobei $h_{in}=c_p \cdot T_{in}$ und $h_{out}=c_p \cdot T_{out}$

Therefore the temperature of air at the outlet side is:

$$T_{out} = T_{in} + \frac{c_{in}^2 - c_{out}^2}{2 \cdot c_p} = (25 + 273.15)\text{K} + \frac{(10\text{m/s})^2 - (150\text{m/s})^2}{2 \cdot 1000\text{J/K} \cdot \text{kg}} = 286,95\text{K}$$

We now calculate the entropy change during this process:

$$\begin{aligned}\Delta s &= c_p \cdot \ln\left(\frac{T_{out}}{T_{in}}\right) - R \cdot \ln\left(\frac{p_{out}}{p_{in}}\right) \\ &= 1\text{J/K} \cdot g \cdot \ln\left(\frac{280.05\text{K}}{300\text{K}}\right) - 0.29\text{J/K} \cdot g \cdot \ln\left(\frac{1\text{ bar}}{1.5\text{ bar}}\right) = 0.049\text{J/K} \cdot g > 0\end{aligned}$$

Hence it is Nozzle.

2)

In order to determine A_{out}/A_{in} , we need additional the law of conservation of mass and the ideal gas law:

$$\begin{aligned}\dot{m}_{in} &= \dot{m}_{out} \Rightarrow \rho_{in} \cdot c_{in} \cdot A_{in} = \rho_{out} \cdot c_{out} \cdot A_{out} \\ \frac{p_{in}}{\rho_{in}} &= R \cdot T_{in} \quad \text{und} \quad \frac{p_{out}}{\rho_{out}} = R \cdot T_{out} \\ \Rightarrow \frac{A_{in}}{A_{out}} &= \left(\frac{p_{out}}{p_{in}}\right) \cdot \left(\frac{T_{in}}{T_{out}}\right) \cdot \left(\frac{c_{out}}{c_{in}}\right) = \left(\frac{0.8\text{bar}}{1\text{ bar}}\right) \cdot \left(\frac{298.15\text{K}}{286.95\text{K}}\right) \cdot \left(\frac{150\text{m/s}}{10\text{m/s}}\right) = 12.47\end{aligned}$$

$A_{in} > A_{out}$, it is also proved that this pipe is a nozzle.